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If it displaces Au + (aq) from solution, then it has a reduction potential smaller than E + Au + (aq) / Au (s) = 1.68V. But if it does not displace Fe3 + (aq) from solution, then its reduction potential is larger than. E + Fe3 + (aq) / Fe2 + (s) = 0.769V. Therefore, 0V < E + < 0.17V.

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Electrochemistry. Strengths of Oxidizing and Reducing Agents • The more positive the Eo redvalue of the half reaction, the greater is the tendency of the reactant of the half reaction to be reduced and therefore to oxidize other species. Zn(s) + Cu 2+(aq) Zn 2+(aq) + Cu(s) The Eo redvalue for Cu.

Chapter 20 Electrochemistry

1. Electrochemistry is the study of the interchange of chemical and electrical energy. A redox (oxidation-reduction) reaction is a reaction in which one or more electrons are transferred. In a galvanic cell, a spontaneous redox reaction occurs which produces an electric current.

CHAPTER SEVENTEEN ELECTROCHEMISTRY

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Answer. Answer: (a) 3. The ionic conductance of Ba 2+ and Cl - are respectively 127 and 76 ohm-1 at infinite dilution. The equivalent conductance of BaCl 2 at infinite dilution will be (A) 139.5 (B) 203 (C) 279 (D) 101.5. Answer. Answer: (a) 4. When heating one end of a metal plate, the other end gets hot because of (A) the resistance of the ...

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