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EXAMPLES 1A (E) A has
a mass number of zero
and an “atomic
number” of 1. Emission
of this electron has the
effect of transforming a
neutron into a proton.
 ${}_{94}^{241}\text{Pu} \rightarrow {}_{95}^{241}\text{Am} + {}_0^{-1}\text{e}$
1B (E) ${}_{58}^{137}\text{Ni}$ has a
mass number of 137
and an atomic number

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of 28. A positron has a
mass

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following problems.

SECTION 25.1

NUCLEAR RADIATION 1.

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mass number and
atomic number of an

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atom that undergoes

beta decay? 2. A

radioisotope of an
element undergoes

alpha particle decay.

How do the atomic

number and mass

number of the particle

change? 3.

SECTION 25.1

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Guide for Chapter 25.

Terms in this set (37)

Neutron Ejection. when

a neutron is emitted

from the nucleus. ${}^1_0\text{n}$.

Particle for Neutron

Ejection. ${}^5_2\text{He} \rightarrow {}^1_0\text{n} +$

${}^4_2\text{He}$.

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CHAPTER 25 What
You'll Learn You will
trace the history of
nuclear chemistry from
dis-covery to
application. You will
identify types of
radioactive decay and
solve decay rate
problems. You will
describe the reac-tions
involved in nuclear
fission and fusion. You
will learn about appli-
cations of nuclear
reactions

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page 1 - 3 out of 25
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25 NUCLEAR

CHEMISTRY PRACTICE
EXAMPLES 1A (E) A has
a mass number of zero
and an “atomic
number” of 1. Emission
of this electron has the
effect of transforming a
neutron into a proton.
 ${}_{0}^{0}\text{e}^{-}$
94 241 95 241 1 0 +
Pu Am 1B (E) 58 Ni has
a mass number of 58

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and an atomic number
of 28.

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- CHAPTER 25

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A nuclear equation is balanced when A) the same elements are found on both sides of the equation. B) the sum of the mass numbers and the sum of the atomic numbers of the particles and atoms are the same on both sides of the

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equation.

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Study a process that
decreases the number
of slow moving
neutrons in a nuclear
reactor; this is
accomplished by using

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control rods made out of material such as cadmium, which absorbs neutrons.

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Section 25.4 Fission and Fusion of Atomic Nuclei In your textbook, read about the process of by which electrical energy is produced in a nuclear power plant. Use the following diagram to

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complete the passage.

D In a nuclear power plant, energy is produced in rea orc by fission reactions that occur in uranium-containing bars called (I)

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EXAMPLES A, 6— has a
mass number of zero
and an
Answer Key

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- (b) numbers of protons and/or neutrons that confer nuclear stability.
- (c) n/p ratios that confer nuclear stability.
- (d) atomic masses that confer nuclear stability.
- (e) atomic masses that

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indicate fissile

isotopes. 2. The actual mass of a ^{37}Cl atom is 36.966 amu. Calculate the mass defect

- (amu/atom) for a ^{37}Cl atom. (a) 0.623 amu
(b) 0.388 amu

Sample Questions - Chapter 26 - Department of Chemistry

A nuclear fuel. A fissionable isotope must be present in large enough

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quantities to sustain a controlled chain reaction. The radioactive isotope is contained in tubes called fuel rods. A moderator. A moderator slows neutrons produced by nuclear reactions so that they can be absorbed by the fuel and cause additional nuclear reactions. A coolant.

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